

## COURSE SYLLABUS

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NOTE: This syllabus is subject to change during the semester. Please check this syllabus on a regular basis for any updates.

**Department** : Physical Sciences  
**Course Title** : General Inorganic Chemistry  
**Section Name** : CHEM\_1311\_1 General Chemistry I  
**Scheduled** : Monday, Tuesday, Wednesday, and Thursday, 09:30a.m. - 11:35a.m.  
**Start Date** : 06/04/2012  
**End Date** : 07/05/2012  
**Modality** : FACE-TO-FACE  
**Credits** : 3

## INSTRUCTOR INFORMATION

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**Name** : Robert Morris  
**OC Email** : rmorris@odessa.edu  
**OC Phone #** : (432) 335-6596

## COURSE DESCRIPTION

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A lecture course designed as a first college-transfer course for students with some background in physical science. This lecture course covers such topics as chemical stoichiometry, atomic structure, bonding, formulas, equations, gas laws, solutions, etc. The student will be involved in reading information or problems and using critical thinking skills and mathematics to organize the information or to arrive at an answer; also requires student writing skills in order to communicate the information acquired in a written format.

## PREREQUISITES/CO REQUISITES

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1. Pass all sections of the COMPASS exam and be eligible to take College Algebra.
2. Co requisite: CHEM 11011 LAB

## [SCANS](#)

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1, 3, 6, 9

## COURSE OBJECTIVES & LEARNING OUTCOMES

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The objective in any chemistry course is to develop problem-solving skills. To find strategies that help you develop the *chemical intuition* needed to understand chemical reasoning.

Upon completion of this course, students will be able to:

- (1) use dimensional analysis with proper attention to units and significant figures
- (2) balance chemical equations and use stoichiometric relationships to quantify reactant and product amounts
- (3) apply Gas Laws

## COURSE ATTENDANCE

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Attendance is required and will be checked. To effectively master the material, your **attendance** for each class is necessary.

### **Required**

Simple Scientific calculator (**Programmable calculators cannot be used on quizzes or exams**)

### **Cell Phone**

When class begins, all phones need to be on silent or vibrate. If it is necessary for you to answer your phone or text message, then you need to step out into the hallway. If you consistently text message while in class, you will be asked to leave the class. Your phone **cannot** be used as a calculator on quizzes or exams.

## **Homework**

Assignments will be made, collected and graded for completeness. The homework assignments are to help you prepare for the exams because variations of these will be on the exam. You should attempt to work all of the homework questions. Answer key will be posted the day before it's due.

## **Quizzes**

There will be a short quiz given once each chapter is completed. The quizzes will include questions and problems over the current material but may require some knowledge of the previous material.

## **Examinations**

(Tentatively)

First Exam:	Thursday Friday, June 15 <sup>th</sup>
Second Exam:	Thursday, July 22 <sup>th</sup>
Third Exam:	Thursday, July 29 <sup>th</sup>
Final Exam:	Thursday, July 5 <sup>th</sup>

**\*NOTE: The due dates are subject to change. Please check this syllabus on a regular basis for any updates.**

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## **REQUIRED READINGS/MATERIALS**

1. *Chemistry*, 12<sup>th</sup>, Brown, 2011.
2. Simple Scientific Calculator

## **SPECIAL NEEDS**

Odessa College complies with Section 504 of the Vocational Rehabilitation Act of 1973 and the Americans with Disabilities Act of 1990. If you have any special needs or issues pertaining to your access to and participation in this or any other class at Odessa College, please feel free to contact me to discuss your concerns. You may also call the Office of Disability services at 432-335-6861 to request assistance and accommodations.

## **LEARNING RESOURCE CENTER (LIBRARY)**

The Library, known as the [Learning Resources Center](#), provides research assistance via the [LRC's catalog \(print books, videos, e-books\)](#) and [databases \(journal and magazine articles\)](#). [Research guides](#) covering specific subject areas, [tutorials](#), and the ["Ask a Librarian"](#) service provide additional help.

## **STUDENT-MAIL**

Please access your [Odessa College Student E-mail](#), by following the link to either set up or update your account: <http://www.odessa.edu/gmail/>. **All assignments or correspondence will be submitted using your Odessa College email.**

## **STUDENT PORTAL**

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## **ODESSA COLLEGE STUDENT SUCCESS COACHES**

The Odessa College Student Success Coaches will help you stay focused and on track to complete your educational goals. If an instructor sees that you might need additional help or success coaching, he or she may submit a Retention Alert or a Starfish Alert. A Student Success Coach will contact you to work toward a solution.

## **TECHNICAL SUPPORT**

Blackboard username and password help may be accessed online course availability and student email account contact the Student Success Center at 432-335-6878 or online at [https://www.odessa.edu/dept/ssc/helpdesk\\_form.htm](https://www.odessa.edu/dept/ssc/helpdesk_form.htm).

## IMPORTANT SCHOOL POLICIES

For information regarding student support services, academic dishonesty, disciplinary actions, special accommodations, or students and instructors' right to academic freedom can be found in the [Odessa College Student Handbook](#).

## COURSE REQUIREMENTS (Lectures, Assignments, Assessments, and Exams)

CHAPTER: Title	DESCRIPTIONS	EVALUATE	QUANTITY	NUMBER
<b>ONE:</b> Introduction: Matter and Measurement	How Science is Done	Quiz	29	1, 2, 4, 7, 8, 10, 12, 14, 17, 19, 20, 27, 31, 33, 35-43, 46, 48, 55, 58, 64, 71
<b>TWO:</b> Atoms, Molecules, and Ions	Atomic Theory	Quiz	31	2, 4, 5, 6, 7, 10, 15, 19, 20, 22, 24, 26, 31, 32, 37, 38, 43, 45, 48, 50, 55, 56, 58, 61, 64, 66, 68, 70, 73, 74, 99
<b>THREE:</b> Stoichiometry: Calculations with Chemical Formulas and Equations	What is a mole?	Quiz	24	1, 4, 5, 6, 7, 11, 12, 14, 16, 18, 22, 24, 34, 36, 38, 46, 50, 54, 62, 68, 77, 84, 93, 103
		<b>Exam I</b>		
<b>FOUR:</b> Reaction in aqueous solution	Types of Reactions	Quiz	25	2, 7, 12, 15, 16, 19, 20, 22, 24, 28, 29, 35, 37, 40, 47, 50, 52, 62, 63, 64, 71, 78, 80, 87, 99
<b>FIVE:</b> Thermochemistry	Enthalpy	Quiz	17	3, 8, 15, 27, 32, 39, 41, 43, 53, 55, 63, 66, 69, 71, 73, 85, 102
<b>SIX:</b> Electronic structure of an atom	Electron Arrangement	Quiz	14	2, 5, 13, 18, 23, 27, 35, 53, 65, 66, 68, 70, 71, 81
		<b>Exam II</b>		
<b>SEVEN:</b> Periodic Properties of the Elements	Trends of Periodic Table	Quiz	13	25, 26, 29, 30, 37, 41, 45, 47, 61, 68, 73, 81, 103
<b>EIGHT:</b> Basic Concepts of Chemical Bonding	Bonding Types	Quiz	16	4, 7, 13, 18, 19, 39, 40, 41, 42, 47, 48, 63, 69, 71, 85, 90
<b>NINE:</b> Molecular Geometry and Bonding Theory	Multiple Bonding	Quiz	17	3, 7, 12, 14, 19, 27, 28, 32, 43, 44, 55, 56, 61, 66, 81, 83, 101
		<b>Exam III</b>		
<b>TEN:</b> Gases	Gas Laws	Quiz	7	4, 10, 21, 35, 37, 43, 53, 54
<b>Final Exam</b>		<b>Exam</b>		

## GRADING POLICY

### Course Grading

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|-----------------------------|--------------------|
| 1. Chapter exams will be:   | 60% of your grade. |
| 3. Chapter Quizzes will be: | 10% of your grade. |
| 4. Homework will be:        | 10% of your grade. |
| 5. Project will be:         | 05% of your grade. |
| 6. Final Exam will be:      | 15% of your grade. |

Percentage %	Grade
>89.5	A
89.4- 79.5	B
79.4- 69.5	C
69.4- 59.5	D
<59.4	F